CHE – LAWS OF CHEMICAL COMBINATION

There are four (4) laws of chemical combination which describe the general features of a chemical change. Although research has shown that these laws are only approximately correct, they still form the basis of chemistry.

The Law of Conservation of Mass (or Matter)

The law of conservation of mass states that matter is neither created nor destroyed during chemical reaction but changes from one form to another.

The Law of Definite Proportions

The law of definite proportions states that all pure samples of a particular chemical compound contain similar elements combined in the same proportion by mass.

The Law of Multiple Proportions

The law of multiple proportions states that if two elements, A and B, combine to form more than one chemical compound, then the various masses of one element, A, which combine separately with a fixed mass of the other element, B, are in sample multiple ratio.

The Law of Reciprocal Proportions

The law of reciprocal proportions states that the masses of several elements, A, B, C, which combine separately with a fixed mass of another element, D, are the same as, or simple multiples of the masses in which A, B, C, themselves combine with one another.